

Chapter 6 Covalent Bonds

Section 6.2 Covalent Bonding**(pages 166–188)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

Reading Strategy (page 166)

Relating Text and Visuals As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

| Molecular Models | |
|-----------------------|-----------------------------------|
| Model | Description |
| Electron-dot diagrams | Dots represent valence electrons. |
| Structural formula | |
| Space-filling | |
| Electron cloud | |

Covalent Bonds (pages 166–187)

- Define a covalent bond. _____

- A molecule is a _____ group of atoms that are joined together by one or more covalent bonds. Circle the correct answer.
negative neutral positive
- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. _____
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
 - one
 - two
 - three

Chapter 6 Chemical Bonding Section 2 Covalent Answer Key

R Barnett



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