

Chemistry - Chapter 9 Assessment Answers (Pages 281-282)

42. a) 2- b) 1+ c) 1- d) 3+

43. a) 2+ b) 2+ c) 3+ d) 1+

44. a) barium ion b) iodide ion c) silver ion d) mercury(II) ion

45. Cyanide,  $\text{CN}^-$  and hydroxide,  $\text{OH}^-$

46. a) hydroxide ion b) lead(IV) ion c) sulfate ion d) oxide ion

47. zero; A compound is electrically neutral.

48. The symbols for the cation and anion are written and the charges are balanced with subscripts. The name of the cation is followed by the name of the anion.

49. Determine the charge of the anion then work the formula backwards to find the charge of the transition metal cation needed to give a net charge of zero for the formula unit.

50. The symbols for the cation and anion are written and the charges are balanced with subscripts. Parentheses are used around the polyatomic ion if a subscript is needed. The name of the cation is followed by the name of the anion.

51. a and b

52. When more than a single polyatomic ion is needed to balance the formula.

53.  $\text{NH}_4\text{NO}_3$ , ammonium nitrate;  $(\text{NH}_4)_2\text{CO}_3$ , ammonium carbonate;  $\text{NH}_4\text{CN}$ , ammonium cyanide;  $(\text{NH}_4)_3\text{PO}_4$ , ammonium phosphate;  $\text{Sn}(\text{NO}_3)_4$ , tin(IV) nitrate;  $\text{Sn}(\text{CO}_3)_2$ , tin(IV) carbonate;  $\text{Sn}(\text{CN})_4$ , tin(IV) cyanide;  $\text{Sn}_2(\text{PO}_4)_4$ , tin(IV) phosphate;  $\text{Fe}(\text{NO}_3)_3$ , iron(III) nitrate;  $\text{Fe}_2(\text{CO}_3)_3$ , iron(III) carbonate;  $\text{Fe}(\text{CN})_3$ , iron(III) cyanide;  $\text{FePO}_4$ , iron(III) phosphate;  $\text{Mg}(\text{NO}_3)_2$ , magnesium nitrate;  $\text{MgCO}_3$ , magnesium carbonate;  $\text{Mg}(\text{CN})_2$ , magnesium cyanide;  $\text{Mg}_3(\text{PO}_4)_2$ , magnesium phosphate

54. two nonmetals

55. a) tri- b) mono- c) di- d) hexa- e) penta- f) tetra-

56. For formulas, write the correct symbols for each element with a subscript corresponding to the prefix before each element in the name. For naming, name each element in the order given. Use the subscript to determine the prefixes before each element in the name. The name ends in -ide.

57. a)  $\text{BCl}_3$  b) dinitrogen pentoxide c)  $\text{N}_2\text{H}_4$  d) carbon tetrachloride

58. a) hydrochloric acid b)  $\text{H}_2\text{SO}_4$  c) nitric acid d)  $\text{HC}_2\text{H}_3\text{O}_2$

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**RS Peters**



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