

## Chapter 11 Nuclear Chemistry

### Review

#### Introduction

**Chemical Energy** is energy stored in the bonds of atoms and molecules. **Nuclear Energy** is stored through the nucleus of atoms changing their their different nuclei. Nuclear energy is produced by **radioactivity**.

### The Nucleus

#### Composition



#### Mass Number

- Nucleus is composed of two main particles, protons and neutrons.
- The number of protons is the atomic number.
- The number of protons and neutrons together is the mass number.

Remember that some elements have different mass numbers. This is because some elements have more or fewer neutrons.

**Radioactivity** is the emission of an element's protons/neutrons electrons, which can be called **alpha**, **beta**, **gamma**, **positron**, **neutron**, **proton**.

Table 11.1: Properties of alpha, beta and gamma particles			
	Alpha	Beta	Gamma
Symbol	$\alpha$	$\beta$	$\gamma$
Charge	+2	-1	0
Mass	4 u	$\frac{1}{1836}$ u	0
Penetrating power	Low	High	Very High
Ionising power	High	Low	Very Low

#### Types of Radioactive Decay

1. **Alpha decay** is the loss of an  $\alpha$ -particle.



2. **Beta decay** is the loss of a  $\beta$ -particle (a high-speed electron) emitted by the nucleus.  ${}_{Z}^{A}\text{X} \longrightarrow {}_{Z+1}^{A}\text{Y}$



3. **Gamma radiation** is the loss of a  $\gamma$ -ray, which is a high-energy  $\gamma$  radiation that is emitted when a radioactive substance emits an  $\alpha$  or  $\beta$  particle.  ${}_{Z}^{A}\text{X} \longrightarrow {}_{Z}^{A}\text{X} + \gamma$

# Chapter 21 Nuclear Chemistry Mixed Review Answers

**JE Gale**



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### **Table of Contents Chapter 21 Nuclear Chemistry Mixed Review Answers**

1. Understanding the eBook Chapter 21 Nuclear Chemistry Mixed Review Answers
  - The Rise of Digital Reading Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 21 Nuclear Chemistry Mixed Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Personalized Recommendations
  - Chapter 21 Nuclear Chemistry Mixed Review Answers User Reviews and Ratings
  - Chapter 21 Nuclear Chemistry Mixed Review Answers and Bestseller Lists

5. Accessing Chapter 21 Nuclear Chemistry Mixed Review Answers Free and Paid eBooks
  - Chapter 21 Nuclear Chemistry Mixed Review Answers Public Domain eBooks
  - Chapter 21 Nuclear Chemistry Mixed Review Answers eBook Subscription Services
  - Chapter 21 Nuclear Chemistry Mixed Review Answers Budget-Friendly Options
6. Navigating Chapter 21 Nuclear Chemistry Mixed Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 21 Nuclear Chemistry Mixed Review Answers Compatibility with Devices
  - Chapter 21 Nuclear Chemistry Mixed Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Highlighting and Note-Taking Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Interactive Elements Chapter 21 Nuclear Chemistry Mixed Review Answers
8. Staying Engaged with Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 21 Nuclear Chemistry Mixed Review Answers
9. Balancing eBooks and Physical Books Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 21 Nuclear Chemistry Mixed Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Setting Reading Goals Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Fact-Checking eBook Content of Chapter 21 Nuclear Chemistry Mixed Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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