

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Solutions

### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about hydration in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is hydration in which the solvent is water. What is hydration?  
**Solution** is the process of surrounding solute particles with solvent particles to form a solution.

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Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## STUDY GUIDE FOR CONTENT MASTERY

## CHAPTER

## 15

## Section 15.1 continued

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the orientation of the water molecules around the sodium ions and chloride ions. Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

19. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

- true** 20. The overall energy change that occurs when a solution forms is called the heat of solution.
- false** 21. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- false** 22. Solvation continues as long as the solubility rate is less than the crystallization rate.
- true** 23. In a saturated solution, solvation and crystallization are in equilibrium.
- true** 24. Additional water can be dissolved in an unsaturated solution.
- true** 25. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Chapter 18 Study Guide For Content Mastery Teacher Edition

**Katrin Zwirgmaier**



**Chapter 18 Study Guide For Content Mastery Teacher Edition:**

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Discover the Magic : Delia Owens' "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens spins a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting.

These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered.

The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

The Secret History is a brilliant and gripping novel that will keep you wondering until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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