

Study Card: Strong vs. Weak Acids and Bases

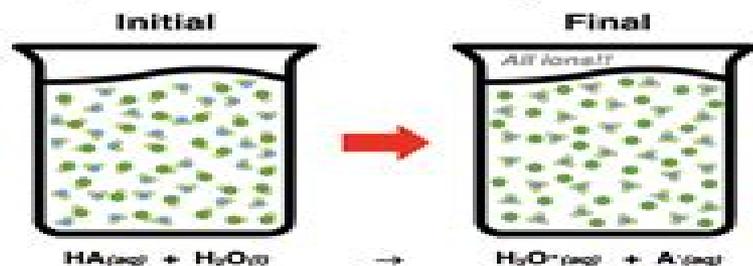
Acids and bases dissociate into ions in aqueous solutions. Whether an acid or a base is classified as strong or weak depends on the **degree of dissociation**.

Strong Acids: completely dissociate into ions when dissolved in water.



Goes to completion.

Equilibrium lies far to the right. $K \gg \gg 1$.

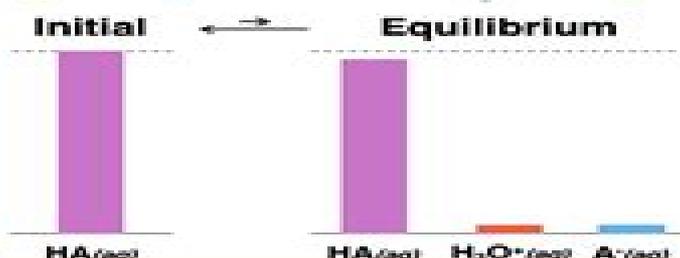
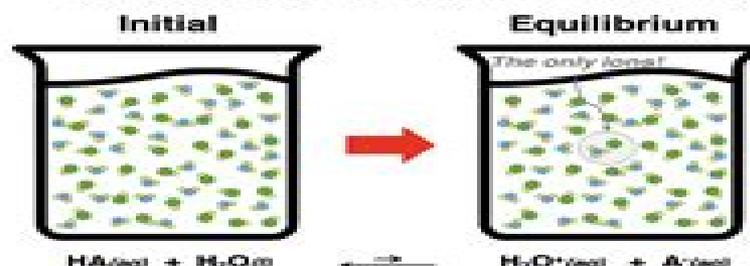


Weak Acids: partially dissociate into ions when dissolved in water.



Reversible; with a strong reverse reaction.

Equilibrium lies far to the left. $K \ll \ll 1$.



Examples

| Strong Acids | Strong Bases |
|---|--|
| Hydrochloric acid (HCl) | Lithium hydroxide (LiOH) |
| Sulfuric acid (H ₂ SO ₄) | Sodium hydroxide (NaOH) |
| Nitric acid (HNO ₃) | Potassium hydroxide (KOH) |
| Hydrobromic acid (HBr) | Rubidium hydroxide (RbOH) |
| Hydroiodic acid (HI) | Cesium hydroxide (CsOH) |
| Perchloric acid (HClO ₄) | Calcium hydroxide (Ca(OH) ₂) |
| Chloric acid (HClO ₃) | Strontium hydroxide (Sr(OH) ₂) |
| | Barium hydroxide (Ba(OH) ₂) |

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33 Strong Vs Weak Acids Answer

AN Whitehead



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A Course of Instruction in the General Principles of Chemistry Arthur Amos Noyes, Miles Standish Sherrill, 1917

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