

## Chapter 9 Practice Problems: Chemical Names and Formulas

### Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
  - oxygen
  - iodine
  - sodium
  - aluminum
  - nickel, two electrons lost
  - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
  - $\text{Cr}^{3+}$
  - $\text{P}^{3-}$
  - $\text{Li}^+$
  - $\text{Ca}^{2+}$
  - $\text{Cl}^-$
  - $\text{O}^{2-}$
- Name each ion. Identify each as a cation or an anion.
  - $\text{Sn}^{2+}$
  - $\text{Co}^{3+}$
  - $\text{Br}^-$
  - $\text{K}^+$
  - $\text{H}^-$
  - $\text{Mn}^{2+}$
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
  - carbonate ion
  - nitrite ion
  - sulfate ion
  - hydroxide ion
  - chromate ion
  - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
  - $\text{CN}^-$
  - $\text{HCO}_3^-$
  - $\text{PO}_4^{3-}$
  - $\text{Cl}^-$
  - $\text{Ca}^{2+}$
  - $\text{SO}_3^{2-}$

### 9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
  - magnesium oxide
  - tin(II) fluoride
  - potassium iodide
  - aluminum chloride
  - sodium sulfide
  - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
  - $\text{Ba}^{2+}$  and  $\text{Cl}^-$
  - $\text{Ag}^+$  and  $\text{I}^-$
  - $\text{Ca}^{2+}$  and  $\text{S}^{2-}$
  - $\text{K}^+$  and  $\text{Br}^-$
  - $\text{Al}^{3+}$  and  $\text{O}^{2-}$
  - $\text{Fe}^{2+}$  and  $\text{O}^{2-}$
- Name the following binary ionic compounds.
  - $\text{MnO}_2$
  - $\text{Li}_3\text{N}$
  - $\text{CaCl}_2$
  - $\text{SrBr}_2$
  - $\text{NiCl}_2$
  - $\text{K}_2\text{S}$
  - $\text{Fe}_3\text{N}_2$
  - $\text{SnCl}_4$
- Write formulas for the following ionic compounds.
  - sodium phosphate
  - magnesium sulfate
  - sodium hydroxide
  - potassium cyanide
  - ammonium chloride
  - potassium dichromate
- Write formulas for the following ionic compounds.
  - $\text{NH}_4^+$  and  $\text{SO}_4^{2-}$
  - $\text{K}^+$  and  $\text{NO}_3^-$
  - barium ion and hydroxide ion
  - lithium ion and carbonate ion

# Chapter 9 Chemical Names And Formulas Practice Problems Answers

**Martha Mackin**



## Chapter 9 Chemical Names And Formulas Practice Problems Answers:

**Barron's Chemistry Practice Plus: 400+ Online Questions and Quick Study Review** Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2022-07-05 Barron's Chemistry Practice Plus features more than 400 online practice questions and a concise review guide that covers the basics of Chemistry. Inside you'll find concise review on the basics of Chemistry, an excellent resource for students who want a quick review of the most important topics, access to 400 online questions arranged by topic for customized practice, online practice includes answer explanations with expert advice for all questions plus scoring to track your progress. This essential guide is the perfect practice supplement for students and teachers.

**Understanding Chemistry** Fred M. Dewey, 1994

**General Organic and Biological Chemistry** Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health-related fields such as nursing, dental hygiene, nutrition, medical technology, and occupational therapy. It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry. Students need have no previous background in chemistry but should possess basic math skills. The text features numerous helpful problems and learning features.

**Addison-Wesley Chemistry**, 1990

**Student's Guide to Introduction to Chemical Principles** by Edward I. Peters, 2d Ed Peter P. Berlow, 1978

*Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross-reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter, and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds, chemical formulas, and the mole, chemical reactions, calculations based on equations, gases, and the properties of a liquid. The remaining chapters examine the solutions, acids, bases, salts, oxidation-reduction reactions, electrochemistry, chemical kinetics, and equilibrium, and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students.

**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

**Scientific Soapmaking** Kevin M. Dunn, 2010 Scientific Soapmaking

bridges the gap between the technical and craft literature It explains the chemistry of fats oils and soaps and teaches sophisticated analytical techniques that can be carried out using equipment and materials familiar to makers of handcrafted soap *The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB **Partial Solutions Manual** Darrell D. Ebbing, David Bookin, 1998

Prentice Hall Chemistry, 2000 **New Technical Books** New York Public Library, 1925 **The Digester, Over the Spillway**, 1976 Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *Chemical Engineering Catalog*, 1925 **Chemistry** James E. Brady, Fred Senese, 2004-02-04 Publisher Description Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction *The Software Encyclopedia 2001*, 2001 **Machinery**, 1907 *R.R. Bowker's Software for Schools*, 1987

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