



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
 - H_2
 - O_2

- Classify each of the following as either ionic or molecular compound. Write the name of the compound.
- K_2SO_4
- H_2O
- H_2SO_4
- H_2O_2
- H_2
- O_2

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. CH_4
- Show covalent bonding in carbon dioxide. CO_2
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules.
 - H_2O
 - H_2O_2
 - H_2S
 - H_2SO_4
 - H_2SO_3
 - H_2SO_2
 - H_2SO
 - H_2S_2
 - $\text{H}_2\text{S}_2\text{O}_7$
 - $\text{H}_2\text{S}_2\text{O}_8$
 - $\text{H}_2\text{S}_2\text{O}_9$
 - $\text{H}_2\text{S}_2\text{O}_{10}$
 - $\text{H}_2\text{S}_2\text{O}_{11}$
 - $\text{H}_2\text{S}_2\text{O}_{12}$
 - $\text{H}_2\text{S}_2\text{O}_{13}$
 - $\text{H}_2\text{S}_2\text{O}_{14}$
 - $\text{H}_2\text{S}_2\text{O}_{15}$
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 - $\text{H}_2\text{S}_2\text{O}_{28}$
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 - $\text{H}_2\text{S}_2\text{O}_{98}$
 - $\text{H}_2\text{S}_2\text{O}_{99}$
 - $\text{H}_2\text{S}_2\text{O}_{100}$

SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by sp^3 ? The shape of the orbital and sub-orbital.
- Which orbital and sub-orbital form the sp^3 hybrid orbital? What is the shape of the orbital and sub-orbital?

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