

Biology Exam #2 (Chapters 2, 3, 4, 6, 7, 8)

Chapter 2:

Atoms are made up of electrons (- outside the nucleus), protons(+inside the nucleus), and neutrons (0 inside the nucleus)

Atomic number= number of protons

Mass number=protons + nucleus

Matter is composed of elements and exists as solid, liquid, gas, or plasma

Element: a substance that cannot be broken down by chemical means to a simpler substance that has different physical properties.

Basic elements to life include C, H, N, O, P, S

Life's diversity is based on carbon, almost all the molecules of a cell are composed of carbon

Organic molecules: carbon-based molecules, molecules have C and H, those with only C and H are called hydrocarbons

Isotopes: atoms of an element with different number of neutrons

Radioactive isotope emits energy as rays and subatomic particles

Uses of radiation include medical imaging, killing bacteria and fungi, killing cancer cells

For neutral atoms, the number of valence electrons is equal to the atom's main group number

ATP: currency of the cell, high energy molecule, undergoes hydrolysis which releases energy

Polar covalent bonds: sharing between two atoms that is unequal

Oxygen is more electronegative than Hydrogen

Second strongest

Nonpolar covalent bonds: sharing of electrons between two atoms is fairly equal

Strongest

Carbon bonds: four electrons on outer shell

Electronegativity: attraction of an atom for electrons in covalent bond

Hydrogen bonding: attraction between negative oxygen and positive hydrogen

Bonds is weak individually, but strong together(weakest)

Biology Study Answers

RJ Shavelson



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Table of Contents Biology Study Answers

1. Understanding the eBook Biology Study Answers
 - The Rise of Digital Reading Biology Study Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Biology Study Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Biology Study Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Biology Study Answers
 - Personalized Recommendations

- Biology Study Answers User Reviews and Ratings
- Biology Study Answers and Bestseller Lists
- 5. Accessing Biology Study Answers Free and Paid eBooks
 - Biology Study Answers Public Domain eBooks
 - Biology Study Answers eBook Subscription Services
 - Biology Study Answers Budget-Friendly Options
- 6. Navigating Biology Study Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Biology Study Answers Compatibility with Devices
 - Biology Study Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Biology Study Answers
 - Highlighting and Note-Taking Biology Study Answers
 - Interactive Elements Biology Study Answers
- 8. Staying Engaged with Biology Study Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Biology Study Answers
- 9. Balancing eBooks and Physical Books Biology Study Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Biology Study Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Biology Study Answers
 - Setting Reading Goals Biology Study Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Biology Study Answers
 - Fact-Checking eBook Content of Biology Study Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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