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# CHAPTER 11

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## EMULSION BREAKING

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An emulsion is an intensive mixture of two liquid phases, such as oil and water, in which the liquids are mutually insoluble and where either phase may be dispersed in the other. In water chemistry, two types of emulsions are commonly found, oily wastewater and emulsified oil water or O/W emulsions and water oil emulsions (water emulsified in oil or W/O emulsions).

Oily water and water oil emulsions can usually be differentiated visually. O/W emulsions appear to be just oily, dirty water; a drop of the emulsion added to water disperses (Figure 11.1a). A W/O emulsion is usually thick and viscous; a drop of this emulsion added to water does not disperse (Figure 11.1b).

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### OIL-IN-WATER EMULSIONS

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An oily water emulsion, in which oil is dispersed in the water phase, may consist any of various types of oil in a wide range of concentrations. These oils are defined as substances that can be extracted from water by hexane, carbon tetrachloride, chloroform, or benzene. In addition to oils, typical contaminants of these emulsions may be solids, silt, silted particles, emulsifiers, floccants, soaps, surfactants, and other solids. The types of oils found in these emulsions will depend on the industry. They may be bita, lubricants, cooking fluids, heavy hydrocarbons such as tars, grease, crude oils, and diesel oils, and also light hydrocarbons including gasoline, kerosene, and jet fuel. Their concentrations in the wastewater may vary from only a few parts per million to as much as 1 to 20% by volume.

A water O/W emulsion is a colloidal system of electrically charged oil droplets surrounded by an ionic environment. Vigorous mixing and shearing of oily wastewater in transfer storage disperses these minute oil droplets throughout the water. Emulsion stability is maintained by a combination of physical and chemical mechanisms. These emulsions act similar in behavior to the colloidal systems encountered in suspensions (settling) and rivers (silt).

One such stabilizing mechanism, emulsions, is brought about by the addition of surface-active agents, such as organic materials or chemicals, which aid in maintaining a stable colloidal system. These molecules usually carry an electric charge and seek out the oil-water interface of the emulsified droplet. Here, the pronounced charges cause the emulsion to be stabilized through repulsion of the surrounding charged droplets. Neutral compounds nevertheless can also stabilize an emulsion, since these molecules are bifunctional: one end is soluble in water, and the other end is hydrocarbon, so the molecule bridges the interface and stabilizes it.

# Chapter 11 Emulsion Breaking Ssu

**Michal Rosen-Zvi**



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## **Table of Contents Chapter 11 Emulsion Breaking Ssu**

1. Understanding the eBook Chapter 11 Emulsion Breaking Ssu
  - The Rise of Digital Reading Chapter 11 Emulsion Breaking Ssu
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 Emulsion Breaking Ssu
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 11 Emulsion Breaking Ssu
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 Emulsion Breaking Ssu

- Personalized Recommendations
- Chapter 11 Emulsion Breaking Ssu User Reviews and Ratings
- Chapter 11 Emulsion Breaking Ssu and Bestseller Lists
- 5. Accessing Chapter 11 Emulsion Breaking Ssu Free and Paid eBooks
  - Chapter 11 Emulsion Breaking Ssu Public Domain eBooks
  - Chapter 11 Emulsion Breaking Ssu eBook Subscription Services
  - Chapter 11 Emulsion Breaking Ssu Budget-Friendly Options
- 6. Navigating Chapter 11 Emulsion Breaking Ssu eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 11 Emulsion Breaking Ssu Compatibility with Devices
  - Chapter 11 Emulsion Breaking Ssu Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 11 Emulsion Breaking Ssu
  - Highlighting and Note-Taking Chapter 11 Emulsion Breaking Ssu
  - Interactive Elements Chapter 11 Emulsion Breaking Ssu
- 8. Staying Engaged with Chapter 11 Emulsion Breaking Ssu
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 11 Emulsion Breaking Ssu
- 9. Balancing eBooks and Physical Books Chapter 11 Emulsion Breaking Ssu
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 11 Emulsion Breaking Ssu
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 11 Emulsion Breaking Ssu
  - Setting Reading Goals Chapter 11 Emulsion Breaking Ssu
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 11 Emulsion Breaking Ssu

- Fact-Checking eBook Content of Chapter 11 Emulsion Breaking Ssu
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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