

Chapter 10 Guided Notes

Name \_\_\_\_\_

1. \_\_\_\_\_ attractions are bonds that hold compounds together
  - a. What are 2 examples of these attractions?
2. \_\_\_\_\_ attractions are between molecules.
3. Which attraction is stronger?
4. Which state of matter will have the **strongest** intermolecular forces? Which state has the **weakest**?



5. Order the intermolecular forces (IMF) in order of weakest to strongest: hydrogen bonding, dipole-dipole, ion-dipole, London dispersion
6. List all the possible IMF for the following compounds:
  - a. HCl
  - b. NH<sub>3</sub>
  - c. H<sub>2</sub>O
  - d. CH<sub>4</sub>
  - e. NaCl dissolved in water

7. For each type of compound determine what forces are possible:

Type of IMF	Atoms	Nongpolar molecules	Polar molecules without OH, NH, or HF groups	Polar molecules containing OH, NH, or HF groups	Ionic solids dissolved in polar liquids
Dispersion forces					
Dipole-dipole					
Hydrogen bonding					
Ion-dipole					

Explore the following procedure and complete the subsequent practice exercises in preparation for the synchronous class discussion. Step 1 or Step 2 are review from general chemistry I. Review Lewis structures

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**B Lingard**



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## **Table of Contents Chapter 10 Guided Practice Problems Answers Pearson**

1. Understanding the eBook Chapter 10 Guided Practice Problems Answers Pearson
  - The Rise of Digital Reading Chapter 10 Guided Practice Problems Answers Pearson
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Guided Practice Problems Answers Pearson
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 10 Guided Practice Problems Answers Pearson
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Guided Practice Problems Answers Pearson
  - Personalized Recommendations
  - Chapter 10 Guided Practice Problems Answers Pearson User Reviews and Ratings
  - Chapter 10 Guided Practice Problems Answers Pearson and Bestseller Lists

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  - Chapter 10 Guided Practice Problems Answers Pearson Public Domain eBooks
  - Chapter 10 Guided Practice Problems Answers Pearson eBook Subscription Services
  - Chapter 10 Guided Practice Problems Answers Pearson Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Chapter 10 Guided Practice Problems Answers Pearson Compatibility with Devices
  - Chapter 10 Guided Practice Problems Answers Pearson Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 10 Guided Practice Problems Answers Pearson
  - Highlighting and Note-Taking Chapter 10 Guided Practice Problems Answers Pearson
  - Interactive Elements Chapter 10 Guided Practice Problems Answers Pearson
8. Staying Engaged with Chapter 10 Guided Practice Problems Answers Pearson
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 10 Guided Practice Problems Answers Pearson
9. Balancing eBooks and Physical Books Chapter 10 Guided Practice Problems Answers Pearson
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 10 Guided Practice Problems Answers Pearson
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 Guided Practice Problems Answers Pearson
  - Setting Reading Goals Chapter 10 Guided Practice Problems Answers Pearson
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 10 Guided Practice Problems Answers Pearson
  - Fact-Checking eBook Content of Chapter 10 Guided Practice Problems Answers Pearson
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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