

Chapter 2: Atomic Structure and Chemical Bonding

- Materials → Molecules → Atoms
 - Atoms = protons (p) + neutrons (n) + electrons (e)
 - Protons and neutrons are made of quarks
 - Quantitative measurements need units: metric or S.I. (Système International) or mks (meter-kilogram-second) units
- | | | |
|-------------|-------------------|-------------------------|
| meter | (m) | for length |
| cubic meter | (m ³) | for volume |
| kilogram | (kg) | for mass |
| Kelvin | (K) | for temperature |
| second | (s) | for time |
| mole | (mol) | for amount of substance |

Chapter 2

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Typical parameters: atoms are small

	Mass (kg)	Charge (coulomb)
Proton	1.673×10^{-27}	$+1.602 \times 10^{-19}$
Neutron	1.675×10^{-27}	0
Electron	9.109×10^{-31}	-1.602×10^{-19}

- Nuclear radius is measured in $10^{-15}\text{m} = \text{fm}$
- Electron radii measured in $\text{Å} = 0.1 \text{ nm} = 10^{-10} \text{ m}$
- Fundamental charge = $e = -1.602 \times 10^{-19} \text{ C}$

Chapter 2

2

Chapter 2 Atomic Structure And Chemical Bonding

Lauren Gardner



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