

17

WATER AND AQUEOUS SYSTEMS

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SECTION 17.1 LIQUID WATER AND ITS PROPERTIES (pages 475–478) ◆◆◆

This section explains the hydrogen bonding that occurs in water molecules and how high surface tension and low vapor pressure can be explained in terms of hydrogen bonding.

► The Water Molecule (pages 475–476)

- What unique molecular compound is the foundation of all life on Earth?
water
- Approximately what fraction of Earth's surface is covered in water? 3/4
- Circle the letter next to each sentence that is true concerning water molecules.
 - Each O — H covalent bond in a water molecule is nonpolar.
 - In a water molecule, the less electronegative hydrogen atoms acquire a slight positive charge and the oxygen atom acquires a slight negative charge.
 - Because the water molecule has an H — O — H bond angle of 105°, the molecule as a whole is polar.
- The diagram below depicts a water molecule. Complete the labels showing the locations of the hydrogen atoms, the oxygen atom, and the regions of positive and negative charge.



- List five important properties of water that result from hydrogen bonding.
 - high surface tension
 - low vapor pressure
 - high specific heat capacity
 - high heat of vaporization
 - high boiling point

Chapter 15 Water And Aqueous Systems Test Answers

Annelies Wilder-Smith



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Table of Contents Chapter 15 Water And Aqueous Systems Test Answers

1. Understanding the eBook Chapter 15 Water And Aqueous Systems Test Answers
 - The Rise of Digital Reading Chapter 15 Water And Aqueous Systems Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 15 Water And Aqueous Systems Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 15 Water And Aqueous Systems Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 15 Water And Aqueous Systems Test Answers
 - Personalized Recommendations
 - Chapter 15 Water And Aqueous Systems Test Answers User Reviews and Ratings
 - Chapter 15 Water And Aqueous Systems Test Answers and Bestseller Lists
5. Accessing Chapter 15 Water And Aqueous Systems Test Answers Free and Paid eBooks
 - Chapter 15 Water And Aqueous Systems Test Answers Public Domain eBooks
 - Chapter 15 Water And Aqueous Systems Test Answers eBook Subscription Services
 - Chapter 15 Water And Aqueous Systems Test Answers Budget-Friendly Options

6. Navigating Chapter 15 Water And Aqueous Systems Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 15 Water And Aqueous Systems Test Answers Compatibility with Devices
 - Chapter 15 Water And Aqueous Systems Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 15 Water And Aqueous Systems Test Answers
 - Highlighting and Note-Taking Chapter 15 Water And Aqueous Systems Test Answers
 - Interactive Elements Chapter 15 Water And Aqueous Systems Test Answers
8. Staying Engaged with Chapter 15 Water And Aqueous Systems Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 15 Water And Aqueous Systems Test Answers
9. Balancing eBooks and Physical Books Chapter 15 Water And Aqueous Systems Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 15 Water And Aqueous Systems Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 15 Water And Aqueous Systems Test Answers
 - Setting Reading Goals Chapter 15 Water And Aqueous Systems Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 15 Water And Aqueous Systems Test Answers
 - Fact-Checking eBook Content of Chapter 15 Water And Aqueous Systems Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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