

VOLUMETRIC ANALYSIS

AIM

Determine the concentration of aqueous hydrogen peroxide and the percentage purity of oxalate in an unknown oxalate sample by titration with potassium permanganate.

INTRODUCTION

Volumetric analysis is a quantitative analysis process whereby volumes are measured. A volume of reagent known as the standard solution of known concentration is chemically reacted with a solution of unknown concentration in order to determine the concentration of the unknown. The solution that is used as standard is called the titrant which is delivered from a burette and is added until the reaction is completed. The end point of the titration is noted by observing the change of colour of an indicator. In a direct titration, the titrant is added to analyte until the end point is observed whereas in a back titration, a known excess of a standard reagent is added to the analyte. A second standard reagent is then used to titrate the excess of the first reagent.

All the calculations in determining the molarity and average concentration for all equations involve the formulae:

$$M_A V_A = a$$

$$M_B V_B = b$$

The equation for the standardization experiment is given in the ionic form as:



The back titration is explained by the equations:



APPARATUS AND REAGENTS

1.6g of Potassium permanganate

Boiling distilled water

Distilled water

Erlenmeyer flask

Rubber stopper

Sodium oxalate

Sulphuric acid

Source of heat

Burette

Hydrogen peroxide

A Volumetric Analysis Lab Report Answers

Michael Brown



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Table of Contents A Volumetric Analysis Lab Report Answers

1. Understanding the eBook A Volumetric Analysis Lab Report Answers
 - The Rise of Digital Reading A Volumetric Analysis Lab Report Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying A Volumetric Analysis Lab Report Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an A Volumetric Analysis Lab Report Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from A Volumetric Analysis Lab Report Answers
 - Personalized Recommendations
 - A Volumetric Analysis Lab Report Answers User Reviews and Ratings
 - A Volumetric Analysis Lab Report Answers and Bestseller Lists
5. Accessing A Volumetric Analysis Lab Report Answers Free and Paid eBooks

- A Volumetric Analysis Lab Report Answers Public Domain eBooks
 - A Volumetric Analysis Lab Report Answers eBook Subscription Services
 - A Volumetric Analysis Lab Report Answers Budget-Friendly Options
6. Navigating A Volumetric Analysis Lab Report Answers eBook Formats
- ePub, PDF, MOBI, and More
 - A Volumetric Analysis Lab Report Answers Compatibility with Devices
 - A Volumetric Analysis Lab Report Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of A Volumetric Analysis Lab Report Answers
 - Highlighting and Note-Taking A Volumetric Analysis Lab Report Answers
 - Interactive Elements A Volumetric Analysis Lab Report Answers
8. Staying Engaged with A Volumetric Analysis Lab Report Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers A Volumetric Analysis Lab Report Answers
9. Balancing eBooks and Physical Books A Volumetric Analysis Lab Report Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection A Volumetric Analysis Lab Report Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine A Volumetric Analysis Lab Report Answers
- Setting Reading Goals A Volumetric Analysis Lab Report Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of A Volumetric Analysis Lab Report Answers
- Fact-Checking eBook Content of A Volumetric Analysis Lab Report Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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