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<u>bug</u>	<u>gum</u>	<u>sun</u>

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Phonics Practice This Week

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+$ is

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3 $\text{Br}^- + \text{BrO}_3^- \rightarrow \text{Br}_2$ previous topic next topic

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- \rightarrow 2\text{H}_2\text{O} + \text{H}_2\text{O}_2$ make electrons equal

convert to basic solution $4\text{cn} + 2\text{au} + 2\text{au} + \text{cn} + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{h}_2\text{o} + \text{o}_2 + \text{h}_2\text{o}_2 + 2\text{oh}^-$ the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{ni} + \text{s} + \text{ni} + 2\text{a} + \text{q} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + \text{Fe}^{2+} + \text{Fe}^{3+} + \text{e}^- = 2$ equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 6\text{Fe}^{2+} + 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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