

Date Performed: 10/26/09

Date Submitted: 11/02/09

Instructor:

Name:

Partner:

## Acid-Base Titrations

### Objective

- To standardize a sodium hydroxide solution (NaOH) with a hydrochloric acid solution (HCl) of known concentration by the reaction.
- To determine the molar mass of an unknown, monoprotic acid
- To determine the percentage of Acetylsalicylic Acid in a tablet of aspirin.

### Experimental Data

#### Part A-Standardization of Sodium Hydroxide Solution

Titration	Vol NaOH (mL)
1	20.00
2	20.21

#### Part B-Molar Mass of an Unknown Monoprotic Acid

Sample	Mass Beaker (g)	Mass Sample + Beaker (g)	Vol NaOH (mL)
1	119.50	120.50	20.95
2	88.32	89.32	19.95

#### Part C-Percentage of Acetylsalicylic Acid in a Commercial Aspirin Tablet

Mass Beaker (g)	Mass Tablet +Beaker (g)	Vol NaOH (mL)
122.08	122.46	8.49

### Sample Calculations

#### 1) Molarity NaOH

(Molarity HCl)(Volume HCl)=(Molarity NaOH)(Volume NaOH)

(0.200M)(25.00mL)=(Molarity NaOH)(20.00mL)

5.00=X(20.00)

Trail 1: Molarity of NaOH=0.250M NaOH

Trail 2: Molarity of NaOH=0.247M NaOH

Average=0.249M NaOH

# Acid Base Titrations Pre Lab Answers Chem Fax

**Yi-Tong Ma**



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