

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make, via the system, visible use of LeChatelier's principle through applying LeChatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
4. In part 4, what is the conjugate acid of the base $\text{Co}(\text{H}_2\text{O})_6^{2+}$ given as $\text{Co}(\text{H}_2\text{O})_6^{2+}$ in the lab? What is the conjugate base of the acid $\text{Co}(\text{H}_2\text{O})_6^{2+}$ given as $\text{Co}(\text{H}_2\text{O})_6^{2+}$ in the lab? What is the conjugate acid of the base $\text{Co}(\text{H}_2\text{O})_6^{2+}$ given as $\text{Co}(\text{H}_2\text{O})_6^{2+}$ in the lab? What is the conjugate base of the acid $\text{Co}(\text{H}_2\text{O})_6^{2+}$ given as $\text{Co}(\text{H}_2\text{O})_6^{2+}$ in the lab?
5. Write the chemical equation for the reaction in part 5.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
6. Write the chemical equation for the reaction in part 6.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
7. Write the chemical equation for the reaction in part 7.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
8. Write the chemical equation for the reaction in part 8.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
9. Write the chemical equation for the reaction in part 9.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
10. Write the chemical equation for the reaction in part 10.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$

Investigating Chemical Equilibrium Lab 12a Answer

Marcel A. Müller



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