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web jun 8 2023 $V = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 1.00 \text{ atm} \times 22.4 \text{ l}$ in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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web aug 3 2022 volumes of gases and liquids are measured in cubic centimetres cm^3 or cubic decimetres dm^3 or litres l 1 mole of any gas at room temperature and pressure occupies 24dm^3 $24\,000\text{cm}^3$ this value is known as the molar volume

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