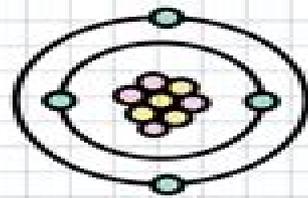


Atomic structure and the periodic table

Atomic structure



protons
 neutrons
 electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

Equations, mixtures and chromatography

chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

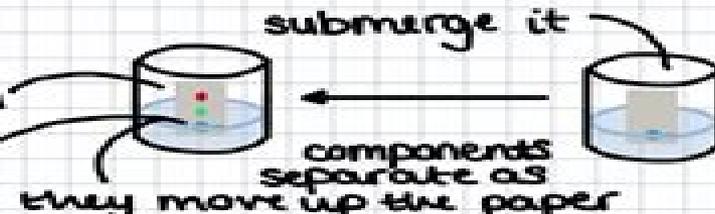
products

word equation: methane + oxygen → carbon dioxide + water

symbol equation: $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

chromatogram

insoluble components stay on the base line



nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

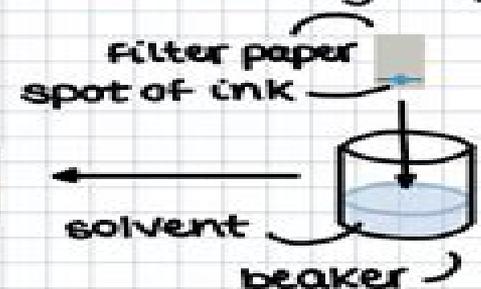
molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

paper chromatography



Atomic Structure Guided Notes Answers

Yan Bai



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Table of Contents Atomic Structure Guided Notes Answers

1. Understanding the eBook Atomic Structure Guided Notes Answers
 - The Rise of Digital Reading Atomic Structure Guided Notes Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Guided Notes Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Guided Notes Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Guided Notes Answers
 - Personalized Recommendations
 - Atomic Structure Guided Notes Answers User Reviews and Ratings
 - Atomic Structure Guided Notes Answers and Bestseller Lists
5. Accessing Atomic Structure Guided Notes Answers Free and Paid eBooks
 - Atomic Structure Guided Notes Answers Public Domain eBooks
 - Atomic Structure Guided Notes Answers eBook Subscription Services
 - Atomic Structure Guided Notes Answers Budget-Friendly Options
6. Navigating Atomic Structure Guided Notes Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Atomic Structure Guided Notes Answers Compatibility with Devices
 - Atomic Structure Guided Notes Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Atomic Structure Guided Notes Answers
 - Highlighting and Note-Taking Atomic Structure Guided Notes Answers
 - Interactive Elements Atomic Structure Guided Notes Answers
 8. Staying Engaged with Atomic Structure Guided Notes Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Guided Notes Answers
 9. Balancing eBooks and Physical Books Atomic Structure Guided Notes Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atomic Structure Guided Notes Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Atomic Structure Guided Notes Answers
 - Setting Reading Goals Atomic Structure Guided Notes Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Atomic Structure Guided Notes Answers
 - Fact-Checking eBook Content of Atomic Structure Guided Notes Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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