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Mark Steyvers



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web nov 9 2021 remove the tajine from the oven set the temperature to broil and using tongs move the chicken from the tajine to a baking sheet leaving behind the sauce olives and preserved lemons broil the chicken skin side up for 5 minutes or until the skin is golden brown and crispy

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web jun 16 2020 christine benlafquih distribute some of the spice mixture over the meat and onions you can use up to 2 3 of the mixture at this step concentrating the seasoning on the onions so the spices will meld with the oil and liquids to make a rich flavorful sauce the reserved spices will be used to season the vegetables

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web 1 consider the reaction $i_2O_5(g) + 5CO(g) \rightarrow 5CO_2(g) + I_2(g)$ a 80.0 grams of iodine v oxide i_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}_2(\text{NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

18g H₂O is your theoretical yield so percent yield = $\frac{149.4\text{g H}_2\text{O}}{180\text{g H}_2\text{O}} \times 100 = 83\%$

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N₂ what is the limiting reagent if 8.52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced

step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ mol O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$2 \text{ mol Al} \times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} = 1 \text{ mol Al}_2\text{O}_3$$

$$1 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g/mol} = 101.964 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield dfrac text actual yield text theoretical yield times 100 percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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80 g i_2O_5 1 mol i_2O_5 1 mol i_2 xs 1 333 8 g i_2O_5 1 mol i_2O_5 28 g co 1 mol co

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to
give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3

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