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**Guillaume Favre**



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web sep 22 2021 the equilibrium system can be written in the general form  $c_e + h_a \rightleftharpoons h_{aq} + a_{aq}$  label 7 the equilibrium constant expression for this reaction is  $K = \frac{c_e \cdot h_{ce}}{c_e \cdot h_a}$  label 8 where we denote the equilibrium constant  $k$  with a subscript  $a$  for acid

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web the study of reaction rates will lead you into the study of chemical equilibrium in this semester you will also study electrochemistry from batteries to electroplating and you will learn about acids bases and salts unit 15 is a brief study of the atomic nucleus which your teacher may introduce at any time during the year

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web equilibrium moles  $\text{hscn}$   $2.00 \times 10^{-5}$  mol  $3.00 \times 10^{-6}$  mol  $1.70 \times 10^{-5}$  mol  $\text{hscn}$  knowing the number of moles of  $\text{Fe}^{3+}$  and  $\text{hscn}$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species mol  $\text{Fe}^{3+}$   $3.170 \times 10^{-5}$  mol  $\text{Fe}^{3+}$   $3.850 \times 10^{-4}$  m

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