



Foldable Phone Prices Returns

Jianjun Gao



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web limiting reactant theoretical and percentage yields key $4KO_2 + 2H_2O \rightarrow 4KOH + 3O_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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web limiting reagents by caroline gumer home using limiting reagents percentage yield answer key percentage yield 1 for the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C_3H_8 what is the percent yield

$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

 8.3 mole $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield = $137.1 \text{ g } \text{H}_2\text{O} / 149.4 \text{ g } \text{H}_2\text{O} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
 if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
 $21.3 \text{ g } \text{O}_2 / 32 \text{ g/mole } \text{O}_2 = 0.666 \text{ moles } \text{O}_2$
 $0.666 \text{ moles } \text{O}_2 \times 3 \text{ g } \text{CO} / \text{mole } \text{CO} = 1.998 \text{ g } \text{CO}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

 b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 / \text{mol } \text{Al}_2\text{O}_3 = 150.0 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{101.96 \text{ g } \text{Al}_2\text{O}_3}{150.0 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 68.0\%$

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web 7. Na₂CO₃ is the limiting reactant percent yield = 86.6% only four molecules can be made 9. this amount cannot be weighed by ordinary balances and is worthless 10. nitrogen is the limiting reagent 11. yes methane is the limiting reagent 12.

c is the limiting reagent 4 33 g of h₂ are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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