

Vol. of 1 mol of ideal gas at STP/NTP = 22.4 L

उसका सही अंश
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoga ans me.

*** Significant fig**
 0 to the left of the 1st non 0 digit are sig. not sig.
 eg. $0.0012 \Rightarrow 2$
 Terminal 0 are not sig. if there is no decimal pt.
 eg. $1000 \Rightarrow 1$ fig.
 0 at end is sig. if there is decimal pt.
 Exponential term are not considered sig.

अंशों - 3/1/31
 $a \times b = c$ } - add/subtract is least amount
 3 same sabse kam fig hoga vo sig. fig ans me hoga

*** LAW OF MASS CONSV**
 In a reactⁿ \rightarrow actual mass of reactant = mass of prod.
 Rasm me zidna use hua.
*** Exceptⁿ \rightarrow Nuclear Rxn** ($E = mc^2$)
 • Pppwedⁿ \rightarrow Balanced Rxn

MOLE CONCEPT

*** LAW OF CONST PROPORTⁿ**
 Elements in a comp are in fixed ratio by mass.
 eg. CO_2
 $12 \quad 32 \text{g} \Rightarrow 3:8$
 Not valid for isomer.
 $\% \text{ of C} = \frac{3}{11} \times 100$

Q. cal. the mass of C req to form 132g of CO_2 .
 $\Rightarrow \text{C} + 2\text{O} \rightarrow \text{CO}_2$
 $12 \text{g} \quad 32 \text{g} \quad 44 \text{g}$
 44g CO_2 ke lie 12g C
 $\therefore 132 \text{g ke lie} = \frac{12}{44} \times 132 = 36 \text{g}$

*** Law of Multiple Propⁿ**
 elements should be same or same fixed mass
 eg. CO CO_2 1:2 ratio
 other eg. N_2O , N_2O_3 , N_2O_5
 exceptⁿ $\rightarrow \text{N}_2\text{O}_3$
 • Average to 140

*** Same (P, T, R, V) of molecules of gases** some (mass eg. 1 mol of)
 $\text{Xe}_2, \text{Xe}_3, \text{Xe}_4$
 $2 : 3 : 4$ molecules
 $2 : 3 : 4$ atoms (different)

*** Gay Lussac's law**
 CONDINS: +T, P same
 • at least 2 gases.
 $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 $1 : 1 : 2$
 • Vol. H_2 lie

*** DALTON'S ATOMIC THEORY**
 • Atoms \rightarrow Neither created nor destroyed

*** MOLE CONCEPT**
 $\frac{A}{Z} \times (P+N)$
 Atomic No. (P)
 • Mass of Proton = $1.67 \times 10^{-27} \text{ kg}$
 • Neutron = $1.67 \times 10^{-27} \text{ kg}$
 • Electron = $9.1 \times 10^{-31} \text{ kg}$
 • $m_p \approx 1837 m_e$

*** Atomic mass unit (amu)**
 $1 \text{ amu} = \frac{1}{12} \times \text{Mass of } ^{12}_6\text{C}$
 $1 \text{ amu} = 1.67 \times 10^{-27} \text{ kg}$
 eg. Mass of $^{12}_6\text{C} = 12 \text{ amu}$
*** Relative Atomic mass**
 • Unitless
 • eg. mass. A.M. of C = 12

*** Molecular Mass**
 • $\text{H}_2\text{SO}_4 = 98 \text{ amu}$
 \rightarrow Mass of 1 molecule of H_2SO_4

*** FORMULA MASS**: for ionic comp.
*** MOLE**
 Amount of substance.
 • 1 mole = No particles (6.022×10^{23})
 $1 \text{ amu} \leq 1.67 \times 10^{-27} \text{ kg}$
 $\text{Na} = 6.022 \times 10^{23}$
 $1 \text{ amu} \times \text{Na} = 1 \text{ g}$

Mole Concept Theory Notes Pdfslibforme

Christian G. Meyer



Mole Concept Theory Notes Pdfslibforme:

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Table of Contents Mole Concept Theory Notes Pdfslibforme

1. Understanding the eBook Mole Concept Theory Notes Pdfslibforme
 - The Rise of Digital Reading Mole Concept Theory Notes Pdfslibforme
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Concept Theory Notes Pdfslibforme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Concept Theory Notes Pdfslibforme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Concept Theory Notes Pdfslibforme
 - Personalized Recommendations
 - Mole Concept Theory Notes Pdfslibforme User Reviews and Ratings

- Mole Concept Theory Notes Pdfslibforme and Bestseller Lists
- 5. Accessing Mole Concept Theory Notes Pdfslibforme Free and Paid eBooks
 - Mole Concept Theory Notes Pdfslibforme Public Domain eBooks
 - Mole Concept Theory Notes Pdfslibforme eBook Subscription Services
 - Mole Concept Theory Notes Pdfslibforme Budget-Friendly Options
- 6. Navigating Mole Concept Theory Notes Pdfslibforme eBook Formats
 - ePub, PDF, MOBI, and More
 - Mole Concept Theory Notes Pdfslibforme Compatibility with Devices
 - Mole Concept Theory Notes Pdfslibforme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Concept Theory Notes Pdfslibforme
 - Highlighting and Note-Taking Mole Concept Theory Notes Pdfslibforme
 - Interactive Elements Mole Concept Theory Notes Pdfslibforme
- 8. Staying Engaged with Mole Concept Theory Notes Pdfslibforme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Concept Theory Notes Pdfslibforme
- 9. Balancing eBooks and Physical Books Mole Concept Theory Notes Pdfslibforme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Concept Theory Notes Pdfslibforme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Concept Theory Notes Pdfslibforme
 - Setting Reading Goals Mole Concept Theory Notes Pdfslibforme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Concept Theory Notes Pdfslibforme
 - Fact-Checking eBook Content of Mole Concept Theory Notes Pdfslibforme
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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