

Question 35



In the early 1900s, Robert Millikan performed a series of experiments to measure the charge on tiny oil droplets suspended in an electric field. By analyzing how the droplets moved under gravity and an applied voltage, Millikan determined the fundamental electric charge (the charge on a single electron). From this charge and additional data, the mass of the electron could be calculated. In this passage, you will examine data from similar oil drop experiments to investigate the mass of the electron.



A fine mist of oil droplets is introduced into a transparent chamber situated between two horizontal metal plates. The space between the plates is observed through a microscope or viewing scope. A DC power supply applies a variable voltage across the plates, creating a uniform electric field. The droplets become charged locally through friction or ionization and can be made to rise, fall more slowly, or remain suspended depending on the balance of forces on them. The time it takes for droplets to move between reference markers and the voltage across the plates are recorded. A cross-sectional view: two parallel horizontal plates, with an oil droplet in between. A microscope is shown to the side for observing the droplet's motion. The top

In Experiment 1, why is the droplet's time to fall measured with the electric field turned off?

- A. To observe how the droplet interacts with the charged plates
- B. To determine the droplet's mass-related properties free from electrical forces
- C. To calculate the exact charge on the droplet
- D. To allow the droplet to pick up additional charge

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Table of Contents Act Practice Latest Setup

1. Understanding the eBook Act Practice Latest Setup
 - The Rise of Digital Reading Act Practice Latest Setup
 - Advantages of eBooks Over Traditional Books
2. Identifying Act Practice Latest Setup
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Act Practice Latest Setup
 - User-Friendly Interface
4. Exploring eBook Recommendations from Act Practice Latest Setup
 - Personalized Recommendations
 - Act Practice Latest Setup User Reviews and Ratings
 - Act Practice Latest Setup and Bestseller Lists
5. Accessing Act Practice Latest Setup Free and Paid eBooks

- Act Practice Latest Setup Public Domain eBooks
 - Act Practice Latest Setup eBook Subscription Services
 - Act Practice Latest Setup Budget-Friendly Options
6. Navigating Act Practice Latest Setup eBook Formats
- ePub, PDF, MOBI, and More
 - Act Practice Latest Setup Compatibility with Devices
 - Act Practice Latest Setup Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Act Practice Latest Setup
 - Highlighting and Note-Taking Act Practice Latest Setup
 - Interactive Elements Act Practice Latest Setup
8. Staying Engaged with Act Practice Latest Setup
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Act Practice Latest Setup
9. Balancing eBooks and Physical Books Act Practice Latest Setup
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Act Practice Latest Setup
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Act Practice Latest Setup
- Setting Reading Goals Act Practice Latest Setup
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Act Practice Latest Setup
- Fact-Checking eBook Content of Act Practice Latest Setup
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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