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$$m = \frac{M \cdot 1000}{1000 - d \cdot M}$$

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web 1) molarity  $M = \frac{m}{V}$  mass molar mass  $\times 0.4500 \text{ L} = 0.825 \text{ g} \times 141.9579 \text{ g/mol} = 0.1172 \text{ mol}$  2) molality  $m = \frac{m}{M}$   $0.825 \text{ g} \times 141.9579 \text{ g/mol} = 0.00581158 \text{ mol}$  3) mole fraction  $n_2 = 0.825 \text{ g} \times 141.9579 \text{ g/mol} = 0.1172 \text{ mol}$   $n_1 = 0.4500 \text{ kg} = 450 \text{ g}$   $n_1 = \frac{450 \text{ g}}{18.015 \text{ g/mol}} = 24.97918401 \text{ mol}$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{mix} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_{total}}$$
 1.5 x 480 + 1.2 x 520 = 480 + 520 = 1000  
 1.344 M with increase in temperature which of these changes a molality b fraction of solute present in water

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