



Lab Data



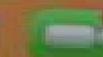
Verify your calculation. In your set-up, do your units cancel?

| | |
|--------------------------------|--------|
| Mass of magnesium (g) | .015 |
| Moles of magnesium (mol) | .00062 |
| Temperature of water (°C) | 23.0 |
| Temperature of water (K) | 296.15 |
| Vapor pressure of water (mmHg) | 21.1 |
| Barometric pressure (mmHg) | 778.0 |

Observations

I observe that there are bubbles that are visible.

| | |
|---|---------|
| Volume of hydrogen gas collected (mL) | 15.00 |
| Volume of hydrogen gas collected (L) | 0.01500 |
| Pressure of hydrogen gas (mmHg) | 756.9 |
| Pressure of hydrogen gas (atm) | 0.9959 |
| Moles of hydrogen gas produced (mol) | .00690 |
| Ideal gas constant (L · atm · mol ⁻¹ · K ⁻¹) | 0.080 |



Introduction To Gas Law Lab Answer Key

David C. Shelledy, Jay I. Peters



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