

GLOBAL  
EDITION



# INTERNATIONAL BUSINESS

Environments & Operations

SIXTEENTH EDITION

John D. Daniels  
Lee H. Radebaugh  
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 Pearson

# International Business Environments And Operations

**M Tight**



## **International Business Environments And Operations:**

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web sep 28 2022 i write down balanced molecular equations for the chemical changes ii write down the number of moles below the formula of each of the reactants and products iii write down the relative masses of the reactants and products with the help of formula below the respective formula these shall be the theoretical amounts of reactants and

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web mass 32 g mass 16 g mass 3 2 g 2 for jars d e f use the number of molecules in the circle to determine the number of moles of o<sub>2</sub> in the jar then calculate the mass of the gas in the jar d e f 12 mol o<sub>2</sub> 2 mol o<sub>2</sub> 6 mol o<sub>2</sub> mass 384 g mass 64 g mass 192 g 3 jars g h and i contain ammonia gas nh<sub>3</sub>

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web a molar mass li<sub>2</sub>co<sub>3</sub> 2 6 941 g 12 01 g 3 16 00 g 73 89 g b molar mass cs<sub>2</sub> 12 01 g 2 32 07 g 76 15 g c molar mass chcl<sub>3</sub> 12 01 g 1 008 g 3 35 45 g 119 37 g d molar mass c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> 6 12 01 g 8 1 008 g 6 16 00 g 176 12 g e molar mass kno<sub>3</sub> 39 10 g 14 01 g 3 16 00 g 101 11 g

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