

Chapter - 1

Chemical Reactions and Equations

Chemical Reaction: • Chemical reaction is a process in which new substances with different properties are formed.

- Atoms of one element don't change into another element.
- Only rearrangement of atoms takes place in a chemical reaction.

Examples 1. Hydrogen + Oxygen \rightarrow water

2. Iron + Oxygen \rightarrow rust

3. Potassium and chlorine gas \rightarrow chloride

Physical Change: • No new substance is formed.

- It is a temporary change.
- No change in the composition of the substance takes place.
- It is easily reversible.
- Heat or light may or may not be given out or absorbed.

Chemical Change: • A new substance is formed.

- It is a permanent change.
- The composition of the new substance change greatly.
- The change may be reversible or irreversible.
- Heat or light or both are generally given out or absorbed.

Law of Conservation of mass

Matter can neither be created nor destroyed in a chemical reaction. This means that total mass of all reactants must

Chemistry Chapter 1 Notes

JG Myers



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one to one property of logarithmic equations one to one property of logarithmic equations for $m > 0$, $n > 0$, $a > 0$ and $a \neq 1$ is any real number if $\log_a m = \log_a n$ then $m = n$

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 $3.5 \times 43x = 2.16 \times 42p = 4.2p = 1.76 \times 2a = 62.3a = 8.22x = 2.23x = 9.63m = 6m = 6.2m = 10.2x = 2x = 2.2x = 11.10.3x = 10x = 1.10$

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