

# Experiment 18-Chemical Reactions Lab Report

Name: \_\_\_\_\_

Section: \_\_\_\_\_

For this lab, you need to balance each equation, and classify the reaction according to the types of reactions listed in the lab handout using C (combination), D (decomposition), Cu (combustion), S (single replacement) and DR (double replacement). Include what you observed below each reaction.

- |  | <u>Type of Reaction</u> |
|--|-------------------------|
| 1. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 (\text{s}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq})$  | _____                   |
| Observations:  |                         |
| 2. $\underline{\hspace{1cm}} \text{CuSO}_4 (\text{H}_2\text{O})_5 + \text{heat} \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{s}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{g})$  | _____                   |
| Observations:  |                         |
| $\underline{\hspace{1cm}} \text{CuSO}_4 (\text{s}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l}) \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{H}_2\text{O})_5 (\text{s})$  | _____                   |
| Observations:  |                         |
| 3. $\underline{\hspace{1cm}} \text{Zn} (\text{s}) + \underline{\hspace{1cm}} \text{HCl} (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{ZnCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2 (\text{g})$   | _____                   |
| Observations:  |                         |
| $\underline{\hspace{1cm}} \text{H}_2 (\text{g}) + \underline{\hspace{1cm}} \text{O}_2 (\text{g}) \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{g})$ The "popping" reaction induced with the match.  | _____                   |
| Observations:  |                         |
| 4. $\underline{\hspace{1cm}} \text{HCl} (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{CO}_2 (\text{g}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ | _____                   |
| Observations:  |                         |
| 5. $\underline{\hspace{1cm}} \text{H}_2\text{O}_2 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{O}_2 (\text{g}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ (the $\text{MnO}_2$ is a catalyst, it does not react)   | _____                   |
| Observations:  |                         |
| 6. $\underline{\hspace{1cm}} \text{HCl} (\text{aq}) + \underline{\hspace{1cm}} \text{NaOH} (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$  | _____                   |
| Observations:  |                         |
| 7. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{CaCO}_3 (\text{s})$   | _____                   |
| Observations:  |                         |
| 8. $\underline{\hspace{1cm}} \text{Zn} (\text{s}) + \underline{\hspace{1cm}} \text{CuSO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{ZnSO}_4 (\text{aq}) + \underline{\hspace{1cm}} \text{Cu} (\text{s})$   | _____                   |
| Observations:  |                         |
| 9. $\underline{\hspace{1cm}} \text{Cu} (\text{s}) + \underline{\hspace{1cm}} \text{ZnSO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{aq}) + \underline{\hspace{1cm}} \text{Zn} (\text{s})$   | _____                   |
| Observations:  |                         |
| 10. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{NaNO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{Ca}(\text{NO}_3)_2 (\text{aq}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq})$   | _____                   |
| Observations:  |                         |

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## **Chemical Reactions Lab Answers:**

**CliffsNotes AP Chemistry** Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

*Core Science Lab Manual with Practical Skills for Class IX V.* K. Sally,Chhaya Srivastava,Goyal Brothers Prakashan,2019-01-01 Goyal Brothers Prakashan **Geology From Experience** E. Kirsten Peters,Larry E. Davis,2000-11-05 Moving away from the observation and vocabulary focus of traditional physical geology lab manuals Peters and Davis s *Geology from Experience* offers experiments that favor hands on involvement and scientific problem solving Students are asked to use geological tools and techniques analyze data from observation experiment and research solve simple equations and make assessments and relevant predictions This approach class tested with great success by the authors gives students a real taste of the scientific experience by revealing the ways geologists actually do their work **Stoichiometry Unit**

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Almir Badnjevic,2017-03-14 This volume presents the proceedings of the International Conference on Medical and Biological  
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    **Scientific American** ,1883      *The Bookseller, Newsdealer and Stationer* ,1899      *Bulletin* Minnesota. University,1905  
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