

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{3 \text{ O}_2} \times \frac{2 \text{ SO}_2}{2 \text{ SO}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ CO}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{326 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} =$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{3 \text{ H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ H}_2\text{SO}_4} =$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{2 \text{ H}_2\text{S}} \times \frac{3 \text{ S}}{1 \text{ mol H}_2\text{S}} = 10.89$$

$$\frac{12.75 \text{ g SO}_2}{64 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ SO}_2} \times \frac{3 \text{ S}}{1 \text{ mol S}} = 19.13$$

Chapter 4 Chemistry Test Answers

Thomas Evangelist



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General Chemistry Darrell D. Ebbing,Larry K. Krannich,Joan I. Senyk,1996

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Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **GED Science** Cambridge, 1993-11

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