

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Solutions

### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about hydration in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is hydration in which the solvent is water. What is hydration?  
**Solution** is the process of surrounding solute particles with solvent particles to form a solution.

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85

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## STUDY GUIDE FOR CONTENT MASTERY

## CHAPTER

## 15

## Section 15.1 continued

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the orientation of the water molecules around the sodium ions and chloride ions. Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atoms) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

19. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

- True    14. The overall energy change that occurs when a solution forms is called the heat of solution.
- False    15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- False    16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- True    17. In a saturated solution, solvation and crystallization are in equilibrium.
- True    18. Additional water can be dissolved in an unsaturated solution.
- True    19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Content Mastery Chapter 12 Answers

**W. James Popham**



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## **Table of Contents Content Mastery Chapter 12 Answers**

1. Understanding the eBook Content Mastery Chapter 12 Answers
  - The Rise of Digital Reading Content Mastery Chapter 12 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Content Mastery Chapter 12 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Content Mastery Chapter 12 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Content Mastery Chapter 12 Answers
  - Personalized Recommendations
  - Content Mastery Chapter 12 Answers User Reviews and Ratings
  - Content Mastery Chapter 12 Answers and Bestseller Lists
5. Accessing Content Mastery Chapter 12 Answers Free and Paid eBooks
  - Content Mastery Chapter 12 Answers Public Domain eBooks
  - Content Mastery Chapter 12 Answers eBook Subscription Services
  - Content Mastery Chapter 12 Answers Budget-Friendly Options

6. Navigating Content Mastery Chapter 12 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Content Mastery Chapter 12 Answers Compatibility with Devices
  - Content Mastery Chapter 12 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Content Mastery Chapter 12 Answers
  - Highlighting and Note-Taking Content Mastery Chapter 12 Answers
  - Interactive Elements Content Mastery Chapter 12 Answers
8. Staying Engaged with Content Mastery Chapter 12 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Content Mastery Chapter 12 Answers
9. Balancing eBooks and Physical Books Content Mastery Chapter 12 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Content Mastery Chapter 12 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Content Mastery Chapter 12 Answers
  - Setting Reading Goals Content Mastery Chapter 12 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Content Mastery Chapter 12 Answers
  - Fact-Checking eBook Content of Content Mastery Chapter 12 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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