

Chapter 17: Chemical Equilibrium

Worksheet: K_c and K_p

1. For the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$

Initial

- 1. $[\text{NO}_2] = 0.100 \text{ M}$ and $[\text{N}_2\text{O}_4] = 0.000 \text{ M}$
- 2. $[\text{NO}_2] = 0.050 \text{ M}$ and $[\text{N}_2\text{O}_4] = 0.075 \text{ M}$
- 3. $[\text{NO}_2] = 0.025 \text{ M}$ and $[\text{N}_2\text{O}_4] = 0.150 \text{ M}$
- 4. $[\text{NO}_2] = 0.0125 \text{ M}$ and $[\text{N}_2\text{O}_4] = 0.300 \text{ M}$
- 5. $[\text{NO}_2] = 0.00625 \text{ M}$ and $[\text{N}_2\text{O}_4] = 0.600 \text{ M}$
- 6. $[\text{NO}_2] = 0.003125 \text{ M}$ and $[\text{N}_2\text{O}_4] = 1.200 \text{ M}$
- 7. $[\text{NO}_2] = 0.0015625 \text{ M}$ and $[\text{N}_2\text{O}_4] = 2.400 \text{ M}$
- 8. $[\text{NO}_2] = 0.00078125 \text{ M}$ and $[\text{N}_2\text{O}_4] = 4.800 \text{ M}$
- 9. $[\text{NO}_2] = 0.000390625 \text{ M}$ and $[\text{N}_2\text{O}_4] = 9.600 \text{ M}$
- 10. $[\text{NO}_2] = 0.0001953125 \text{ M}$ and $[\text{N}_2\text{O}_4] = 19.200 \text{ M}$

Answers

- 1. $K_c = 10.0$
- 2. $K_c = 10.0$
- 3. $K_c = 10.0$
- 4. $K_c = 10.0$
- 5. $K_c = 10.0$
- 6. $K_c = 10.0$
- 7. $K_c = 10.0$
- 8. $K_c = 10.0$
- 9. $K_c = 10.0$
- 10. $K_c = 10.0$

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