

Chemistry 3 & 4, Chapter 4 – Volumetric Analysis TEST SOLUTIONS

Multiple Choice

Identify the choice that best completes the statement or answers the question.

Chemistry 3 & 4, Chapter 2

Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. For a substance to be suitable as a primary standard in a volumetric analysis, it should have:
- A a known formula with a low molar mass.
 - B a very low melting point
 - C no tendency to absorb moisture from the air
 - D no acid–base or redox properties.

ANSWER: C

POINTS: 0 / 1

FEEDBACK: In most volumetric analyses, the primary standard must be an acid, a base, an oxidant or a reductant. Since it must be a solid at room temperature so that it can be accurately weighed out, it needs to have a relatively high MP. Since it must be accurately weighed out, it should not absorb any substance from the air. But a low molar mass means that there will be a higher percentage of error should the measured mass differ from the true mass by, for example, 0.005 g.

REF: pp48–9

2. In a volumetric analysis of vinegar to determine the concentration of ethanoic acid present, a pair of students obtained the following results:

Initial burette reading (mL)	0.10	0.15	0.00	0.25	0.30
Final burette reading (mL)	22.70	22.60	22.50	22.73	22.60

The mean titre is:

- A 22.60 mL
- B 22.48 mL
- C 22.477 mL
- D 22.47 mL

ANSWER: B

POINTS: 0 / 1

FEEDBACK: The titres are 22.60, 22.45, 22.50, 22.48 and 22.30 mL respectively. The concordant results are 22.45, 22.48 and 22.50 mL. The mean of these is 22.48 mL, since a mean cannot have more decimal places than the results used to calculate it.

REF: pp45–6

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1. Understanding the eBook Multiple Choice Questions Volumetric Analysis
 - The Rise of Digital Reading Multiple Choice Questions Volumetric Analysis
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions Volumetric Analysis
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Questions Volumetric Analysis
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions Volumetric Analysis
 - Personalized Recommendations
 - Multiple Choice Questions Volumetric Analysis User Reviews and Ratings
 - Multiple Choice Questions Volumetric Analysis and Bestseller Lists
5. Accessing Multiple Choice Questions Volumetric Analysis Free and Paid eBooks
 - Multiple Choice Questions Volumetric Analysis Public Domain eBooks
 - Multiple Choice Questions Volumetric Analysis eBook Subscription Services
 - Multiple Choice Questions Volumetric Analysis Budget-Friendly Options

6. Navigating Multiple Choice Questions Volumetric Analysis eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Questions Volumetric Analysis Compatibility with Devices
 - Multiple Choice Questions Volumetric Analysis Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Questions Volumetric Analysis
 - Highlighting and Note-Taking Multiple Choice Questions Volumetric Analysis
 - Interactive Elements Multiple Choice Questions Volumetric Analysis
8. Staying Engaged with Multiple Choice Questions Volumetric Analysis
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Questions Volumetric Analysis
9. Balancing eBooks and Physical Books Multiple Choice Questions Volumetric Analysis
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Questions Volumetric Analysis
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions Volumetric Analysis
 - Setting Reading Goals Multiple Choice Questions Volumetric Analysis
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions Volumetric Analysis
 - Fact-Checking eBook Content of Multiple Choice Questions Volumetric Analysis
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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