

Name \_\_\_\_\_

Molecular Geometries Lab Report Sheet

| Formula                      | Lewis Structure   | No of electron domains/Atp | Electron domain geometry     | VSEPR<br>Molecular geometry | Hybridization of central atom | Polar ? |
|------------------------------|---|----------------------------|------------------------------|-----------------------------|-------------------------------|---------|
| CO <sub>2</sub>              | <p>Valence 24</p> $\text{O}=\text{C}=\text{O}$ <p>2 charge clouds</p> | 0                          | Linear<br>180°               | Linear<br>                  | sp                            | NP      |
| BF <sub>3</sub>              | <p>Valence 32</p> $\text{F}=\text{B}-\text{F}$                        | 0                          | Trigonal planar<br>120°      |                             | sp <sup>2</sup>               | NP      |
| SnCl <sub>2</sub>            | <p>Valence 24</p> $\text{Cl}=\text{Sn}-\text{Cl}$                     | 3                          | Bent<br>120°                 |                             | sp <sup>2</sup>               | P       |
| BF <sub>4</sub> <sup>-</sup> | <p>Valence 40</p> $\left[ \text{F}-\text{B}-\text{F} \right]^{-}$     | 0                          | Tetrahedral<br>109.5°        |                             | sp <sup>3</sup>               | P       |
| AsH <sub>3</sub>             | <p>Valence 24</p> $\text{H}-\text{As}-\text{H}$                       | 4                          | Trigonal pyramidal<br>109.5° |                             | sp <sup>3</sup>               | P       |
| H <sub>2</sub> S             | <p>Valence 14</p> $\text{H}-\text{S}-\text{H}$                        | 2                          | Bent<br>109.5°               |                             | sp <sup>3</sup>               | NP      |
| NH <sub>2</sub> <sup>-</sup> | <p>Valence 10</p> $\left[ \text{H}-\text{N}-\text{H} \right]^{-}$     | 2                          | Bent<br>109.5°               |                             | sp <sup>3</sup>               | P       |
| PCl <sub>4</sub>             | <p>Valence 40</p> $\text{F}-\text{P}-\text{Cl}$                       | 0                          | Trigonal bipyre<br>120 + 90  |                             | sp <sup>3</sup> d             | P       |

# Molecular Geometry Lab Report Answers

**James M. Postma, Julian L. Robert, J.  
Leland Hollenberg**



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