



# Pumpkin Spice Today Tutorial

**Conrad Riker**



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molality  $\frac{\text{moles solute}}{\text{kilograms solvent}}$   
mathematical manipulation of molality is the  
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web 1 molarity  $M$  mass molar mass  $\times 0.4500 \text{ l} = 0.825 \text{ g}$  141 9579  $\text{g mol}^{-1} \times 0.0129 \text{ m} = 2$  molality  $0.825 \text{ g} / 141.9579 \text{ g mol}^{-1} = 0.00581158 \text{ mol kg}^{-1}$   
0.00581158 mol / 0.4500 kg = 0.0129 m  
3 mole fraction  $n_2 / n_{\text{total}} = 0.825 \text{ g} / 141.9579 \text{ g mol}^{-1} / (0.825 \text{ g} / 141.9579 \text{ g mol}^{-1} + 0.00581158 \text{ mol} / 0.4500 \text{ kg}) = 0.18015$   
4 0.825 g / 141.9579 g mol<sup>-1</sup> = 0.00581158 mol  
h 2 o 450.0 g / 18.015 g mol<sup>-1</sup> = 24.97918401 mol

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it is written like this  
 $\text{begin array l} \frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100 \text{ end array}$

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