

PHYSICAL PHARMACY

PHYSICAL CHEMISTRY – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

PHYSICAL PHARMACY – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

I. FORCES OF ATTRACTION

A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na: $p^+ = 11$; $e^- = 11 \rightarrow p^+ = 11$; $e^- = 10 \rightarrow$ charge = +1)

○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex. $H^{\delta+}-Br^{\delta-}$)
- **Non Polar** – equal sharing of electrons (ex. $Br-Br$)

B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole** → P + P; Keesom Forces; orientation effect
- **Dipole-Induced Dipole** → P + NP; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole** → NP + NP; London Forces; dispersive effect

- **Ion-Induced Dipole** → ion + NP

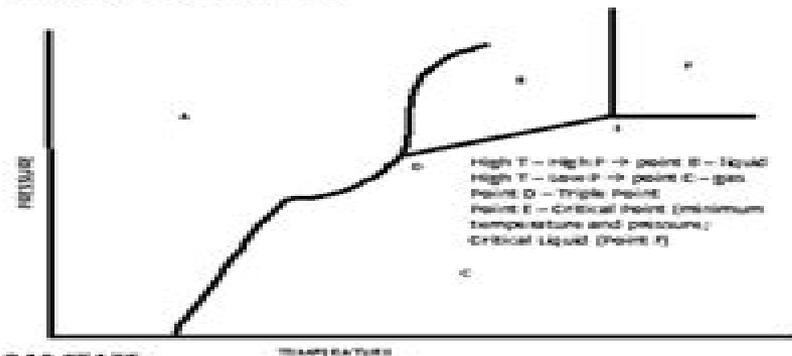
$I_2 \rightarrow NP$ (solubility = 1:2960)

$I_2 + NaI/KI$ (solubilizing agent) → $I_3 + I_2$ = increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



PHASE DIAGRAM



GAS STATE:

- **Kinetic Molecular Theory**

- Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
- Particles do not attract one another but move with complete independence
- Particles exhibit continuous random motion owing to their kinetic energy
- Gases exhibit perfect elasticity (no energy exchange)

II. GAS LAWS

- at standard temperature (Kelvin = $^{\circ}C + 273$) and pressure (1 atm = 760 mmHg or torr; 760 mmHg; 1.01325×10^5 dynes/cm²; 1.01325×10^5 Pa or N/m²; 0.1325 bar)

- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- * density = increase altitude; decrease volume; $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- * H_2O – 1.0g/ml → $4^{\circ}C$

- **Gay Lussac's Law** → constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- * $121^{\circ}C$ – 15-20psi

Physical Pharmacy Lecture Notes

AW Rasmussen



Physical Pharmacy Lecture Notes:

The book delves into Physical Pharmacy Lecture Notes. Physical Pharmacy Lecture Notes is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Physical Pharmacy Lecture Notes, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Physical Pharmacy Lecture Notes
- Chapter 2: Essential Elements of Physical Pharmacy Lecture Notes
- Chapter 3: Physical Pharmacy Lecture Notes in Everyday Life
- Chapter 4: Physical Pharmacy Lecture Notes in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Physical Pharmacy Lecture Notes. This chapter will explore what Physical Pharmacy Lecture Notes is, why Physical Pharmacy Lecture Notes is vital, and how to effectively learn about Physical Pharmacy Lecture Notes.

3. In chapter 2, the author will delve into the foundational concepts of Physical Pharmacy Lecture Notes. The second chapter will elucidate the essential principles that must be understood to grasp Physical Pharmacy Lecture Notes in its entirety.

4. In chapter 3, this book will examine the practical applications of Physical Pharmacy Lecture Notes in daily life. This chapter will showcase real-world examples of how Physical Pharmacy Lecture Notes can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Physical Pharmacy Lecture Notes in specific contexts. The fourth chapter will explore how Physical Pharmacy Lecture Notes is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Physical Pharmacy Lecture Notes. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Physical Pharmacy Lecture Notes.

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