

Grade 11 Physical Sciences Practical: Intermolecular Forces

CAPS & NSC Aligned | 100 Minutes | 25 Learners in 5 Groups

Teacher Version

This practical investigation allows learners to explore the relationship between intermolecular forces (IMF) and physical properties: evaporation (volatility), surface tension, and miscibility. Learners rotate through three stations, collect data, and complete an individual write-up worth 40 marks.

Timing: 100 min

- 0–10 min: Introduction, safety briefing, station overview
- 10–70 min: Station rotations (20 min per station)
- 70–75 min: Pack-up
- 75–100 min: Individual write-up

Grouping: 25 learners → 5 groups of 5

Materials Required

Liquid	Per Group	Class Total (5 Groups)
Water (pure)	20 mL	100 mL
Water + detergent (pre-mix)	10 mL	50 mL
Ethanol (methylated spirits)	20 mL	100 mL
Propan-2-ol (isopropyl)	5 mL	25 mL
Glycerol	5 mL	25 mL
Cooking oil	20 mL	100 mL

Safety Precautions

- Wear goggles and gloves at all times.
- No flames — ethanol and propan-2-ol are flammable.
- Do not ingest or touch chemicals with bare hands.
- Wipe spills immediately to avoid slips.
- Dispose of mixed liquids into labelled waste beakers — follow school policy.

Stations

Station A — Evaporation (Volatility) vs IMF

Liquids: water, ethanol, propan-2-ol, glycerol. Measure 0.50 mL (or 10 drops) into watch glasses. Record time to dryness. Repeat for one liquid.

Station B — Surface Tension (Drops-on-a-Coin) vs IMF

Liquids: water, water + detergent, ethanol. Drop from 2 cm height onto a coin until it spills. Record maximum drops. Repeat for water.

Station C — Miscibility vs IMF & Polarity

Liquids: water (with food colouring), cooking oil, ethanol. Test water + oil, then add ethanol; also test ethanol + oil. Record miscibility and layer positions.

Physical Science Paper 2 Intermolecular Forces Investigation

Hussin A.Rothana



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Table of Contents Physical Science Paper 2 Intermolecular Forces Investigation

1. Understanding the eBook Physical Science Paper 2 Intermolecular Forces Investigation
 - The Rise of Digital Reading Physical Science Paper 2 Intermolecular Forces Investigation
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper 2 Intermolecular Forces Investigation
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Paper 2 Intermolecular Forces Investigation
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper 2 Intermolecular Forces Investigation
 - Personalized Recommendations
 - Physical Science Paper 2 Intermolecular Forces Investigation User Reviews and Ratings
 - Physical Science Paper 2 Intermolecular Forces Investigation and Bestseller Lists
5. Accessing Physical Science Paper 2 Intermolecular Forces Investigation Free and Paid eBooks

- Physical Science Paper 2 Intermolecular Forces Investigation Public Domain eBooks
 - Physical Science Paper 2 Intermolecular Forces Investigation eBook Subscription Services
 - Physical Science Paper 2 Intermolecular Forces Investigation Budget-Friendly Options
6. Navigating Physical Science Paper 2 Intermolecular Forces Investigation eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Paper 2 Intermolecular Forces Investigation Compatibility with Devices
 - Physical Science Paper 2 Intermolecular Forces Investigation Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Paper 2 Intermolecular Forces Investigation
 - Highlighting and Note-Taking Physical Science Paper 2 Intermolecular Forces Investigation
 - Interactive Elements Physical Science Paper 2 Intermolecular Forces Investigation
 8. Staying Engaged with Physical Science Paper 2 Intermolecular Forces Investigation
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Paper 2 Intermolecular Forces Investigation
 9. Balancing eBooks and Physical Books Physical Science Paper 2 Intermolecular Forces Investigation
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Paper 2 Intermolecular Forces Investigation
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Science Paper 2 Intermolecular Forces Investigation
 - Setting Reading Goals Physical Science Paper 2 Intermolecular Forces Investigation
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science Paper 2 Intermolecular Forces Investigation
 - Fact-Checking eBook Content of Physical Science Paper 2 Intermolecular Forces Investigation
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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