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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$   $\Delta h = 2 \text{ h}_2 \text{ (g)} + \text{C}_2\text{H}_6 \text{ (g)}$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)}$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)}$   $\Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ (g)}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{O}_2 \text{ (g)} \rightarrow 2 \text{CO}_2 \text{ (g)} + 2 \text{H}_2\text{O (l)}$   $\Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6 \text{ (g)} + 7/2 \text{O}_2 \text{ (g)} \rightarrow 2 \text{CO}_2 \text{ (g)} + 3 \text{H}_2\text{O (l)}$   $\Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2 \text{ (g)} + 1/2 \text{O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$   $\Delta h = 285 \text{ kJ/mole}$

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changes of its individual steps hess s law example 12 1 determine delta h for the target reaction  $\text{ce } 2 \text{ no}_2 \text{ g } 1 \text{ 2 o}_2 \text{ g } \text{n}_2\text{o}_5 \text{ g}$  given the following information

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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