

The Hydrogen Atom – Student Guide

Background Material

Carefully read the background pages entitled Energy Levels, Light, and Transitions and answer the following questions to check your understanding.

Question 1: Complete the following table which compares how the Bohr Model and the Quantum model represent the Hydrogen atom. In some cases they both describe things in the same way and in some cases they do not.

Concept	Bohr Model	Quantum Model
Nucleus	positively charged proton at the center	positively charged proton at the center
Electrons	Negatively charged cloud.	negatively charged cloud
Orbital	Electrons travel in a circular orbits around nucleus.	Electrons are in a certain position in orbit.
Location of electron in the first orbital	always located at the Bohr radius	Always located on electron cloud.
Energy of the Energy Levels	$E = -E_0/n^2$	$E = -E_0/n^2$

Question 2: Complete the following table which relates the parameters of two different photons by circling the appropriate relationship. The first row is completed for you: "A red photon has a larger wavelength, smaller frequency, smaller energy, and the same velocity through space as a blue photon".

Answers To The Hydrogen Atom Student Guide

Annelies Wilder-Smith



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