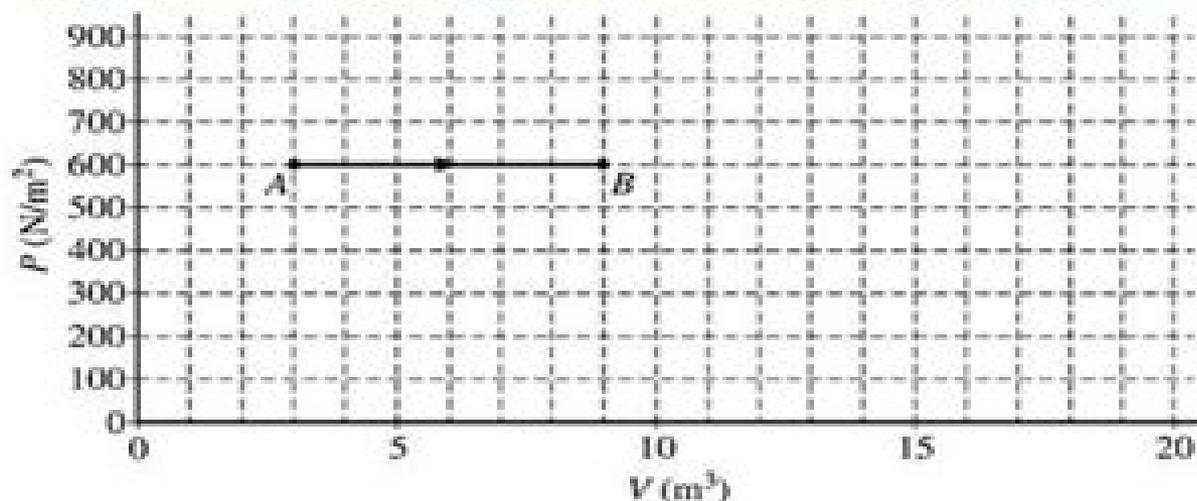


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5. (10 points)

The diagram above of pressure P versus volume V shows the expansion of 2.0 moles of a monatomic ideal gas from state A to state B . As shown in the diagram, $P_A = P_B = 600 \text{ N/m}^2$, $V_A = 3.0 \text{ m}^3$, and $V_B = 9.0 \text{ m}^3$.

(a)

- i. Calculate the work done by the gas as it expands.
- ii. Calculate the change in internal energy of the gas as it expands.
- iii. Calculate the heat added to or removed from the gas during this expansion.

(b) The pressure is then reduced to 200 N/m^2 without changing the volume as the gas is taken from state B to state C . Label state C on the diagram and draw a line or curve to represent the process from state B to state C .

(c) The gas is then compressed isothermally back to state A .

- i. Draw a line or curve on the diagram to represent this process.
- ii. Is heat added to or removed from the gas during this isothermal compression?

_____ added to _____ removed from

Justify your answer.

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