

1 • Matter and Measurement

2 *100 BP* *1F* *212* *of degrees are smaller than °C*
0 FP *32*
PRACTICE TEST

- Which physical state of matter exhibits the greatest change in volume with changes in temperature or pressure? *mostly empty space*
 a) solid b) liquid **c) gas**
easy to compress
- According to the Kinetic Molecular Theory particles of a solid
 a) are bound in a regular array and do not move.
 b) float freely within an array occupying various positions relative to neighbors. *True*
 c) have no relationship to the microscopic structure of the solid.
d) vibrate back and forth but do not move past immediate neighbors.
 e) float freely in the inside but do not move on the surface.
- What volume of a liquid having a density of 1.48 g/cm³ is needed to supply 5.00 grams of the liquid?
 a) 0.296 cm³ **d) 3.38 cm³**
 b) 1.48 cm³ e) 7.40 cm³
 c) 2.26 cm³ *5.00g * 1 cm³ / 1.48g =*
- The density of aluminum is 2.70 g/cm³. If a cube of aluminum weighs 13.5 grams, what is the length of the edge of the cube?
 a) 5.00 cm d) 0.312 cm
b) 1.71 cm e) 0.200 cm
 c) 1.25 cm
*13.5g * 1 cm³ / 2.70g = 5.00 cm³*
∛5.00 cm =

- Which temperature change is the smallest?
 a) 10°C to 20°C **d) 10°F to 20°C**
~~b) 10 K to 20°C~~ **e) 10°F to 20°F**
 c) 10 K to 20 K
- The number of significant figures in 0.06060 x 10⁻⁵ is
 a) 2 d) 5
 b) 3 e) 6
c) 4
- The number, three hundred and fifty thousand, written in scientific notation is best written as
 a) 350 d) 3.50 x 10⁵
 b) 3.5 x 10⁶ e) 3.50 x 10⁻⁵
c) 3.5 x 10⁵ *assume least # of digits significant digits*
350,000 = 3.5 x 10⁵
- The mass of a sample weighted on an electronic balance that is sensitive to ±0.3 mg is 1.2300 g. The number of significant figures in this measurement is *.3mg = .0003g*
 a) 1 d) 4
 b) 2 **c) 5** *1.2300 ± .0003*
 e) 3
- What is the numerical value of:
 1.5 cm - 7.222 x 10⁻¹ cm?
 a) 0.7778 cm **d) 0.8 cm**
 b) 0.778 cm e) 7.072 x 10⁻¹ cm
 c) 0.78 cm

$$\begin{array}{r} 1.5 \text{ cm} \\ - 7.222 \text{ cm} \\ \hline .7778 \text{ cm} \\ \hline .8 \text{ cm} \end{array}$$

Ap Chemistry Chapter 1 Test Wikispaces

John Moore, Richard H. Langley



Ap Chemistry Chapter 1 Test Wikispaces:

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