Specific Heat and Heat Capacity **Name _____KEY_____ Period _____**

1. **Calculate the amount of heat needed to increase the temperature of 250 g of water from 20 °C to 46 °C.**

$q = m \times C \times \Delta T$
$q = 250\ g \times 4.18\ J/g\ °C \times 26°C$
$q = 27,170\ J$ or $27\ kJ$

2. **Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15 g of copper from 25 °C to 60. °C.**

$q = m \times C \times \Delta T$
$C = q/m \times \Delta T$
$C = 204.75\ J / (15\ g \times 35\ °C)$
$C = 0.39\ J/g\ °C$

3. **216 J of energy is required to raise the temperature of aluminum from 15 °C to 35 °C. Calculate the mass of aluminum. (Specific heat capacity of aluminum is 0.90 J/g °C).**

$q = m \times C \times \Delta T$
$m = q/C \times \Delta T$
$m = 216\ J / (0.90\ J/g\ °C \times 20\ °C)$
$m = 12\ g$

4. **The initial temperature of 150 g of ethanol was 22 °C. What will be the final temperature of the ethanol if 3240 J was needed to raise the temperature of the ethanol? (Specific heat capacity of ethanol is 2.44 J/ g °C).**

$q = m \times C \times \Delta T$
$\Delta T = q/m \times C$
$\Delta T = 3240\ J / (150\ g \times 2.44\ J/g\ °C)$
$\Delta T = 8.85\ °C$
$T_{final} = 22\ °C + 8.85\ °C = 31\ °C$

5. **When 435 J of heat is added to 3.4 g of olive oil at 21 °C, the temperature increases to 85 °C. What is the specific heat of the olive oil?**

$q = m \times C \times \Delta T$
$C = q/m \times \Delta T$
$C = 435\ J / (3.4\ g \times 64\ °C)$
$C = 2.0\ J/g\ °C$

6. **How many kilojoules of heat are absorbed when 1.00 L of water is heated from 18 °C to 85 °C? (Hint: You first need to determine the mass of the water, then calculate q in the requested unit.)**

Section 3 Reinforcement Using Heat Answers

R Sandford

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